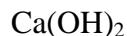


Homework #8
(Naming Acids; Ionic Bonds & Compounds)
CHEM 11 (S 2016, LAC)
(Due Monday, April 18)

1. Name the following acids and bases.



2. Write the formulas for the following compounds

hypochlorous acid--

nickel(II) hydroxide

aluminum hydroxide

hydroiodic acid--

carbonic acid--

3. Write electron configurations for each of the following atoms. Then write the symbol for the most common ion each would form and the electron configuration of that ion.

Cl

Se

Ba

K

4. Use the Born-Haber cycle to determine the net energy from the formation of copper (II) oxide. According to your calculations, will this be a stable compound? Begin by writing the formula for this compound as an end product of the reaction of solid copper metal and molecular oxygen gas. [Notes: IE for Cu = (746, 1958, 3555, etc.) kJ/mol; EA for Cu = 119 kJ/mol; atomization energy of Cu = 339 kJ/mol; lattice energy for copper (II) oxide = 4104 kJ/mol; and the data for oxygen was given in the example of sodium oxide in class.]